

# Limiting reactant questions and answers pdf

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All the revision you need in one place For more practice problems and video lessons visit t mass in grams of  $\text{TiCl}_4$  could be made from g  $\text{TiO}_2$ , g C, and g  $\text{Cl}_3$  is called the “limiting reagent” mol  $\text{NH}_3$ , mol O • In limiting reagent questions we use the limiting reagent as the “given quantity” and. (a) Write a balanced equation for this reaction. Limiting reagents in Key Objectives EXPLAIN how the amount of product in a reaction is affected by an insufficient quantity of any of the reactants The limiting reactant in a reaction is [A] the reactant for which there is the most amount in grams [B] the reactant for which there is the least amount in grams [C] the reactant for which there is the fewest number of moles [D] the reactant which has the lowest coefficient in a balanced equation [E] none of these The limiting reactant is the Answer Sheet 1) Consider the following reaction  $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow (\text{NH}_4)_3\text{PO}_4 + \text{NaNO}_3$  Answer the questions above, assuming we started with grams of ammonium nitrate and grams of sodium phosphate.  $2\text{Al} (\text{s}) + 6\text{HCl} (\text{aq}) \rightarrow 2\text{AlCl}_3 (\text{aq}) + 3\text{H}_2 (\text{g})$  Start with either reactant and solve for the mass of the other reactant; € g Al x 1 mol Al / g Al x 6 mol HCl / 2 mol Al x g HCl / 1 mol HCl = g HCl, we do not have g of HCl Problem For the combustion of sucrose:  $\text{C}_{12}\text{H}_{22}\text{O}_{11} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ . Which is the limiting reagent? What is the limiting reactant? g  $\text{C}_6\text{H}_{12}\text{O}_6$  is = Limiting Reactants worksheets, questions and revision for GCSE Combined Science and Chemistry. Solution path) Calculate moles of sucrose (g/mol = mol) Calculate moles of oxygen required to react with moles of sucrose See answer (b) What mass of pure Ca must be added to excess water to produce g  $\text{C}_2\text{H}_2$ ? there are g of sucrose and g of oxygen reacting. ammonium nitrate is limiting grams of ammonium phosphate, grams of sodium nitrate grams of sodium phosphate is in excess and S is limiting mol S give  $\times = \text{mol S Cl mass} \times = \text{g S Cl}$  this needs  $\times = \text{mol Cl remaining Cl} = \text{mol}$ , mass g Phosphorus trichloride reacts with water according to the stoichiometry:  $\text{PCl}_3 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_3 + \text{HCl}$  A g sample of  $\text{PCl}_3$  g of each reactant is used.

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Étape 1 -

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