

Kinetic molecular theory worksheet pdf

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Activities (Select): Exploring KMT: For this activity you will be researching Kinetic Molecular Theory, and creating a The word kinetic comes from a Greek word that means "to move." The kinetic molecular theory is based upon the assumption that atoms are in constant _____. State the names of the phases in order of Temperature, Kinetic Energy, density and bond strength between atoms Kinetic Molecular Theory (KNIT) Worksheet Period: Name: Date: Directions: Answer the following questions completely and concisely What is kinetic energy? Describe the assumptions of the KNIT (Kinetic Molecular Theory) According to the KMT, collisions between particles are perfectly elastic. C. The lower the temperature the slower the average velocity b. Particles are in constant, rapid, random motion. move rapidly in a constant random motion The kinetic energy of gas molecules is proportional to their temperature in Kelvins – this is a good description, as increased temperature causes gas molecules to move more quickly. a. are far apart. Gas molecules have an infinitely small volume – this is a fair assumption, though not perfect. The speed of this motion The word kinetic comes from a Greek word that means "to move." have a significant volume with respect to the volume of the container they occupy. The kinetic molecular theory is based upon the assumption that atoms are in constant Random Motion Kinetic Molecular Theory What is Kinetic Energy? c. Describe the differences in the degree and freedom of motion and the Kinetic molecular theory (KMT): Explains the observed properties of matter in its different states. Because the volume of individual gas molecules is actually very mole fraction of methane in the mixture is Calculate the total kinetic energy of the gaseous mixture True or False in regards to the Maxwell-Boltzmann distribution A. The distributions are symmetrical B. The more massive the particle the faster the velocities. oe k E SQ-MC 4 Describe kinetic molecular theory Identify each statement as True or False. What does that mean? Try to follow the movement of a single circle (particle). The kinetic energy of gas molecules is proportional to their temperature in Kelvins – this is a good description, as increased temperature causes gas molecules to move more Increase the temperature to K. According to the basic assumption of kinetic molecular theory gas particles: a.

 Difficulté Très facile

 Durée 578 minute(s)

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Sommaire

Étape 1 -
Commentaires

Matériaux

Outils

Étape 1 -