

Aromaticity of benzene pdf

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ChAromatic Compounds, 2 Structure and Unique Properties of Benzene C₆H₆ B C Resonance Structures Facts to Accommodate elements of unsaturation All C-C bonds are same length, not alternating (contrary to expectation based on structure A) 3 • IR: Aromatic ring C-H stretching at cm^{-1} and peaks to cm^{-1} UV: Peak near nm and a less intense peak in nm range 1H NMR: Aromatic H's strongly deshielded by ring and absorb between δ and δ - Peak pattern is characteristic of positions of substituents Spectroscopy of Aromatic Compounds OCHNH benzene toluene phenol aniline styrene. In reaction, the Naming aromatic compounds: (arenes) large number on non-systematic names (Table) CHOH. Generally, mono-substituted benzenes are named in a similar manner as hydrocarbons with -benzene as the parent name. Br Benzene & Aromaticity Chapter Organic Lecture Series Concept of Aromaticity The underlying criteria for aromaticity were recognized in the early s by Erich Hückel, based on molecular orbital (MO) calculations To be aromatic, a compound must be cyclic have one p orbital on each atom of the ring uninterrupted cyclic cloud of π electrons above and below the plane of the ring. The German Chemist Erich Hückel was the first one to recognize that an aromatic compound must have an odd number of pairs of electrons, which can mathematically be written as $4n+2$ ($n = 0, 1, 2, 3$ etc) 1 Chem Jasperse ChNotes. Aromaticity Electron delocalization and resonance Aromatic, antiaromatic, homoaromatic and non-aromatic compounds The reason for this low pKa is its high Benzene & Aromaticity Chapter Organic Lecture Series Concept of Aromaticity The underlying criteria for aromaticity were recognized in the early s by Erich Hückel, Organic Chemistry Chapter Benzene & Aromaticity. Aromaticity. Aromatic rings are very stable. I. BENZENE. Benzene is a six-carbon cyclic structure with alternating single and double bonds • IR: Aromatic ring C-H stretching at cm^{-1} and peaks to cm^{-1} UV: Peak near nm and a less intense peak in nm range 1H NMR: Aromatic H's Aromatic compounds are planar, cyclic systems of conjugation and contains $4n+2^*$ electrons where n is any integer. Lecture Outline.

 Difficulté Difficile

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